

A Four-Coordinate Fe(III) Porphyrin Cation

Ming Fang, Scott R. Wilson, and Kenneth S. Suslick*

School of Chemical Sciences, University of Illinois at Urbana–Champaign, Urbana, Illinois 61801

Received October 21, 2007; E-mail: ksuslick@uiuc.edu

Few families of compounds have been as intensely studied as iron porphyrin complexes. In spite of this, a truly four-coordinate Fe(III) porphyrin cation has never before been reported.¹ The exceptional electrophilicity of the four-coordinate $[\text{Fe}^{\text{III}}(\text{Porph})]^+$ cation will even coordinate arene solvent upon crystallization if the counteranion is extremely non-coordinating, as in the work of Reed with silver dihexabromocborane,^{2,3} and in solution, a Br atom of the hexabromocborane anion binds to the Fe.⁴

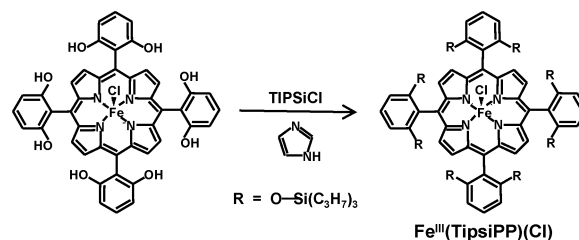
Here we present the first purely four-coordinate Fe(III) porphyrin both in the solid state and in solution. To do so, we have combined both steric and electronic factors by creating a very sterically hindered bis-pocket siloxyl porphyrin in conjunction with a bulky and very weakly coordinating anion.

We have synthesized an extremely hindered bis-pocket siloxyl porphyrin, 5,10,15,20-tetrakis(2',6'-bis(triisopropylsiloxy)phenyl)porphyrin ($\text{H}_2\text{TipsiPP}$) by the reaction of the octahydroxyl iron porphyrin⁵ (5,10,15,20-tetrakis(2',6'-dihydroxyphenyl)porphyrin with triisopropyl chlorosilane using imidazole as a catalyst (Scheme 1). In spite of the driving force of strong Si–O bond formation, high temperatures (250 °C) are required to overcome steric hindrance and fully silylate all eight sites. The resulting porphyrin is extremely sterically hindered with a pocket opening of only 2 Å. Figure 1 compares the space-filling models of the simple $\text{FeTPP}(\text{Cl})$ and the bis-pocket siloxylporphyrin $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{Cl})$. The steric hindrance prevents anion coordination by reducing the thermodynamic binding constant.

To replace the sterically undemanding chloro ligand, a dichloromethane solution of $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{Cl})$ was mixed with 1 equiv of $\text{AgCB}_{11}\text{H}_6\text{Br}_6$ or AgCF_3SO_3 , which produces a color change from yellow to red (Figure 2). NMR, MALDI-TOF, and elemental analysis were used to confirm the purity of the final products. Both complexes react with water readily to form a five-coordinate spin-admixed Fe(III) complex, which has diagnostic UV–vis and NMR features (Figures S1 and S2 in Supporting Information (SI)).

Proton NMR is a very sensitive probe for the spin state of iron.^{4,6,7} On the basis of the β -pyrrole proton NMR isomer shift, Reed has established a magnetochemical series for different anions.⁸ Surprisingly, solutions of $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+$ with a variety of weakly coordinating anions (i.e., $\text{CB}_{11}\text{H}_6\text{Br}_6^-$, SbF_6^- , ClO_4^- , and CF_3SO_3^-) all have exactly the same porphyrin ^1H NMR chemical shifts with β -pyrrole proton chemical shift at –81 ppm in CD_2Cl_2 at 290 K (Figure S3 in SI), which corresponds to the Fe(III) intermediate spin state ($S = 3/2$).⁸ Because these anions are very different both in nucleophilicity and size, the porphyrin NMR spectra would be different if the anions were bound to the iron. This is the case for $\text{Fe}^{\text{III}}\text{TPP}^+$: β -pyrrole proton chemical shifts for $\text{FeTPP}(\text{X})$ are –60 ppm for $\text{CB}_{11}\text{H}_6\text{Br}_6^-$, –31.5 ppm for SbF_6^- , 13 ppm for ClO_4^- , and 39.3 ppm for CF_3SO_3^- .^{8,9} Possible assignments to either $\text{Fe}^{\text{III}}(\text{TipsiPP}^+)$ porphyrin cation radicals or $\text{Fe}^{\text{II}}(\text{TipsiPP})$ were excluded by independent synthesis of the pure complexes, which show very different chemical shifts and have different spin states.

Scheme 1. Synthesis of a Bis-pocket Siloxylporphyrin, $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{Cl})$



The solvent plays a central role in stabilization of the four-coordinate Fe(III) cation. The four-coordinate species is only stable in halogenated solvents (i.e., CH_2Cl_2 , CHCl_3 , CH_2Br_2). In aromatic solvents, such as toluene and benzene, it converts to an admixed spin state, probably due to arene coordination to the iron center.³ The β -pyrrole chemical shift of $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+$ is the same in either CH_2Cl_2 or CH_2Br_2 , which shows that the halocarbons do not perturb the iron center. The compound is not soluble in aliphatic solvents.

Red-brown needle-shaped crystals of $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ were grown by slow evaporation of a 1:1 mixture of dichloro-

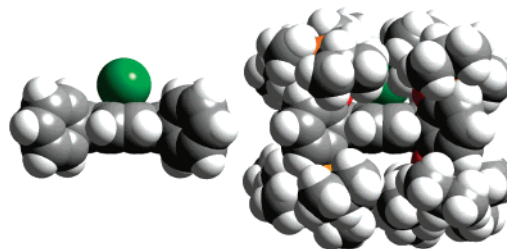


Figure 1. Space-filling model of $\text{Fe}^{\text{III}}\text{TPP}(\text{Cl})$ (left) and $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{Cl})$ (right); side view, same scale.

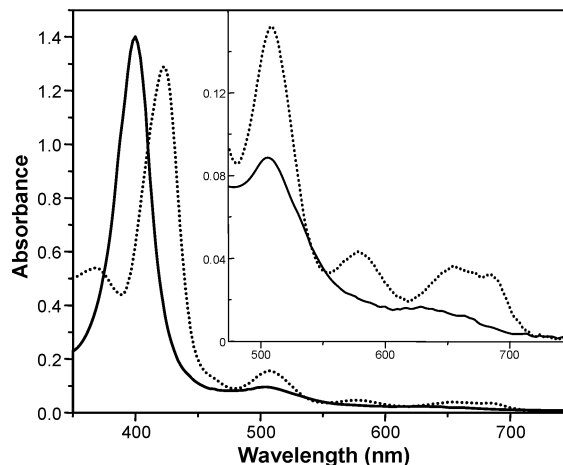


Figure 2. UV–visible spectrum of five-coordinate starting material $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{Cl})$ (dotted line) and four-coordinate $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ (solid line). Inset shows an expanded absorbance scale.

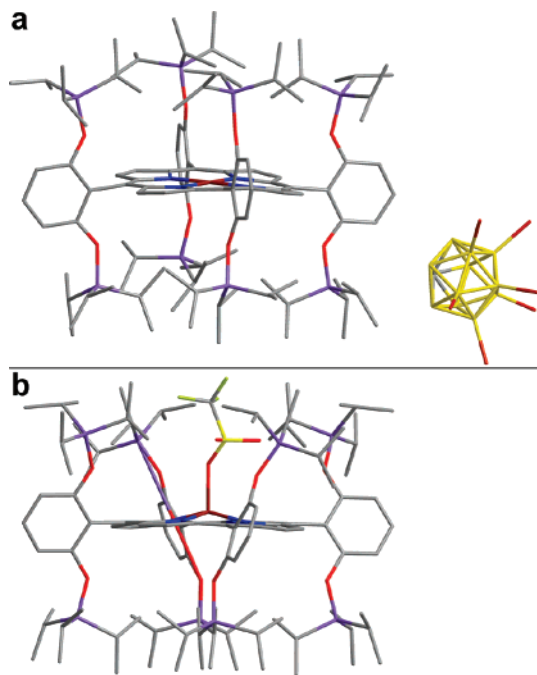


Figure 3. X-ray single-crystal structure of (a) $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ and (b) $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{CF}_3\text{SO}_3)$. Figure S9 in SI provides a crystal packing diagram for $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ showing a layered salt structure with alternating planes of positive and negative charges.

methane/heptane solution inside an inert atmosphere box. The single crystals were redissolved in dry CD_2Cl_2 and tested by NMR to confirm no contamination with water during the crystal growing process. The crystal structure (Figure 3) confirms the absence of axial ligation to the iron porphyrin. The average Fe–N bond length is $1.94 \pm 0.01 \text{ \AA}$ with the iron sitting in the porphyrin plane. This very short bond length is consistent with the removal of the unpaired electron from the $d_{x^2-y^2}$ orbital.^{1,10} The porphyrin is slightly ruffled, with *meso*-carbons above and below the average nitrogen plane by 0.23 \AA . The carborane anion is *definitively not coordinated* to the iron and is fully outside of the porphyrin pocket. The resulting salt has a layered structure with alternating planes of positive and negative charges (a crystal packing diagram is given as Figure S9 in SI). A similar packing structure has been reported for $[\text{FeTPP}][\text{Ag}(\text{CB}_{11}\text{H}_6\text{Br}_6)_2] \cdot 4(p\text{-xylene})$ with the exception of the coordination of the xylene to the iron center.^{2,3}

In the solid state, the four-coordinate iron porphyrin cation can only be isolated with hexabromocarborane as the counterion. In crystals of solids with other anions (e.g., CF_3SO_3^-), the anion enters the porphyrin pocket and coordinates to the iron, presumably driven by the loss of solvation of the anion. The powder X-band EPR of $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{CF}_3\text{SO}_3)$ shows a $g_{\perp}=5.6$ signal, consistent with an admixed spin assignment (Figure S4 in SI). The average Fe–N bond length in $\text{Fe}^{\text{III}}(\text{TipsiPP})(\text{CF}_3\text{SO}_3)$ is $2.053(2) \text{ \AA}$ (which is 0.1 \AA longer than in the true four-coordinate structure), and the iron atom is out of the plane of the four nitrogens by 0.406 \AA . To accommodate the triflate ion, the phenyl groups rotate $\sim 12^\circ$ about the plane normal to porphyrin, which further demonstrates the interplay of electronic and steric factors in these highly hindered systems.

Further characterization data on the four-coordinate $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ are all consistent with a $^3/2$ intermediate spin ground state. The solution magnetic susceptibility of the four-coordinate complex is $4.1 \mu_{\text{B}}$ in a dichloromethane solution

measured by Evan's method from room temperature to 190 K (Figure S5 in SI); SQUID data show no change in the magnetic moment between 300 and 10 K . EPR data of both frozen halocarbon solutions and the crystalline $[\text{Fe}^{\text{III}}(\text{TipsiPP})]^+[\text{CB}_{11}\text{H}_6\text{Br}_6]^-$ have a $g_{\perp} = 4.2$ signal (Figure S4 in SI). Mössbauer spectra were obtained with the ^{57}Fe -enriched sample at 6 K , with an isomer shift of 0.33 mm/s (a typical value for $\text{Fe}(\text{III})$), but with a very large quadrupole splitting value of 5.16 mm/s (Figure S6 in SI). Solid-state NMR also gives an upfield signal at -80 ppm for the β -pyrrole hydrogens, confirming that the four-coordinate $\text{Fe}(\text{III})$ porphyrin is stable both in solid and solution phases.

Although four-coordinate $\text{Fe}(\text{III})$ heme is unlikely in any heme protein, the intermediate spin state of $\text{Fe}(\text{III})$ porphyrins has an interesting counterpart in nature.^{11–17} In cytochrome *c'*, the heme is believed to be a quantum admixed $S = 5/2$ and $S = 3/2$ spin state.^{18–20} Depending on the field strength of the ligand coordinated to the iron, the ratio of $S = 5/2$ to $S = 3/2$ character varies: as the ligand strength becomes weaker, the spin state of iron approaches the pure intermediate spin state of $3/2$.

In summary, the first four-coordinate $\text{Fe}(\text{III})$ porphyrin complex has been isolated and fully characterized in both solution and solid state and shows a pure $S = 3/2$ intermediate spin state. The combination of steric hindrance of a bis-pocket porphyrin with weakly coordinating anions has proved essential in isolation of this highly elusive species, whose reactivity we are still exploring.

Acknowledgment. We gratefully acknowledge the assistance of Alexandre Anastácio and Prof. J. W. Stucki (Mössbauer), Prof. C. M. Rienstra, and Dr. Donghua Zhou (SSNMR), Dr. M. J. Nilges (EPR), and Jin Ho Bang (SQUID). These studies were supported by National Science Foundation (CHE03-15494).

Supporting Information Available: Synthesis, characterization, and X-ray crystallographic details in CIF format. This material is available free of charge via the Internet at <http://pubs.acs.org>.

References

- (1) Scheidt, W. R. In *Porphyrin Handbook*; Kadish, K. M., Smith, K. M., Guillard, R., Eds.; Academic Press: San Diego, CA, 2000; Vol. 3, pp 49–112.
- (2) Evans, D. R.; Fackler, N. L. P.; Xie, Z.; Rickard, C. E. F.; Boyd, P. D. W.; Reed, C. A. *J. Am. Chem. Soc.* **1999**, *121*, 8466–8474.
- (3) Xie, Z.; Bau, R.; Reed, C. A. *Angew. Chem.* **1994**, *106*, 2566–2568; see also *Angew. Chem., Int. Ed. Engl.* **1994**, *2533* (2523/2524), 2433–2564.
- (4) Evans, D. R.; Reed, C. A. *J. Am. Chem. Soc.* **2000**, *122*, 4660–4667.
- (5) Sen, A.; Suslick, K. S. *J. Am. Chem. Soc.* **2000**, *122*, 11565–11566.
- (6) Walker, F. A. *Inorg. Chem.* **2003**, *42*, 4526–4544.
- (7) Walker, F. A. In *Porphyrin Handbook*; Kadish, K. M., Smith, K. M., Guillard, R., Eds.; Academic Press: San Diego, CA, 2000; Vol. 5, pp 81–183.
- (8) Reed, C. A.; Guiset, F. *J. Am. Chem. Soc.* **1996**, *118*, 3281–3282.
- (9) Boersma, A. D.; Goff, H. M. *Inorg. Chem.* **1982**, *21*, 581–586.
- (10) Scheidt, W. R.; Reed, C. A. *Chem. Rev.* **1981**, *81*, 543–555.
- (11) Dolphin, D. H.; Sams, J. R.; Tsin, T. B. *Inorg. Chem.* **1977**, *16*, 711–713.
- (12) Summerville, D. A.; Cohen, I. A.; Hatano, K.; Scheidt, W. R. *Inorg. Chem.* **1978**, *17*, 2906–2910.
- (13) Reed, C. A.; Mashiko, T.; Bentley, S. P.; Kastner, M. E.; Scheidt, W. R.; Spartalian, K.; Lang, G. *J. Am. Chem. Soc.* **1979**, *101*, 2948–2958.
- (14) Goff, H.; Shimomura, E. *J. Am. Chem. Soc.* **1980**, *102*, 31–37.
- (15) Balch, A. L.; Cheng, R. J.; La Mar, G. N.; Latos-Grazynski, L. *Inorg. Chem.* **1985**, *24*, 2651–2656.
- (16) Schunemann, V.; Gerdan, M.; Trautwein, A. X.; Haoudi, N.; Mandon, D.; Fischer, J.; Weiss, R.; Tabard, A.; Guillard, R. *Angew. Chem., Int. Ed.* **1999**, *38*, 3181–3183.
- (17) Nasset, M. J. M.; Cai, S.; Shokhireva, T. K.; Shokhirev, N. V.; Jacobson, S. E.; Jayaraj, K.; Gold, A.; Walker, F. A. *Inorg. Chem.* **2000**, *39*, 532–540.
- (18) Maltempo, M. M.; Moss, T. H. *Q. Rev. Biophys.* **1976**, *9*, 181–215.
- (19) Zeng, Y.; Caignan, G. A.; Bunce, R. A.; Rodriguez, J. C.; Wilks, A.; Rivero, M. J. *Am. Chem. Soc.* **2005**, *127*, 9794–9807.
- (20) Weiss, R.; Gold, A.; Terner, J. *Chem. Rev.* **2006**, *106*, 2550–2579.

JA078061L